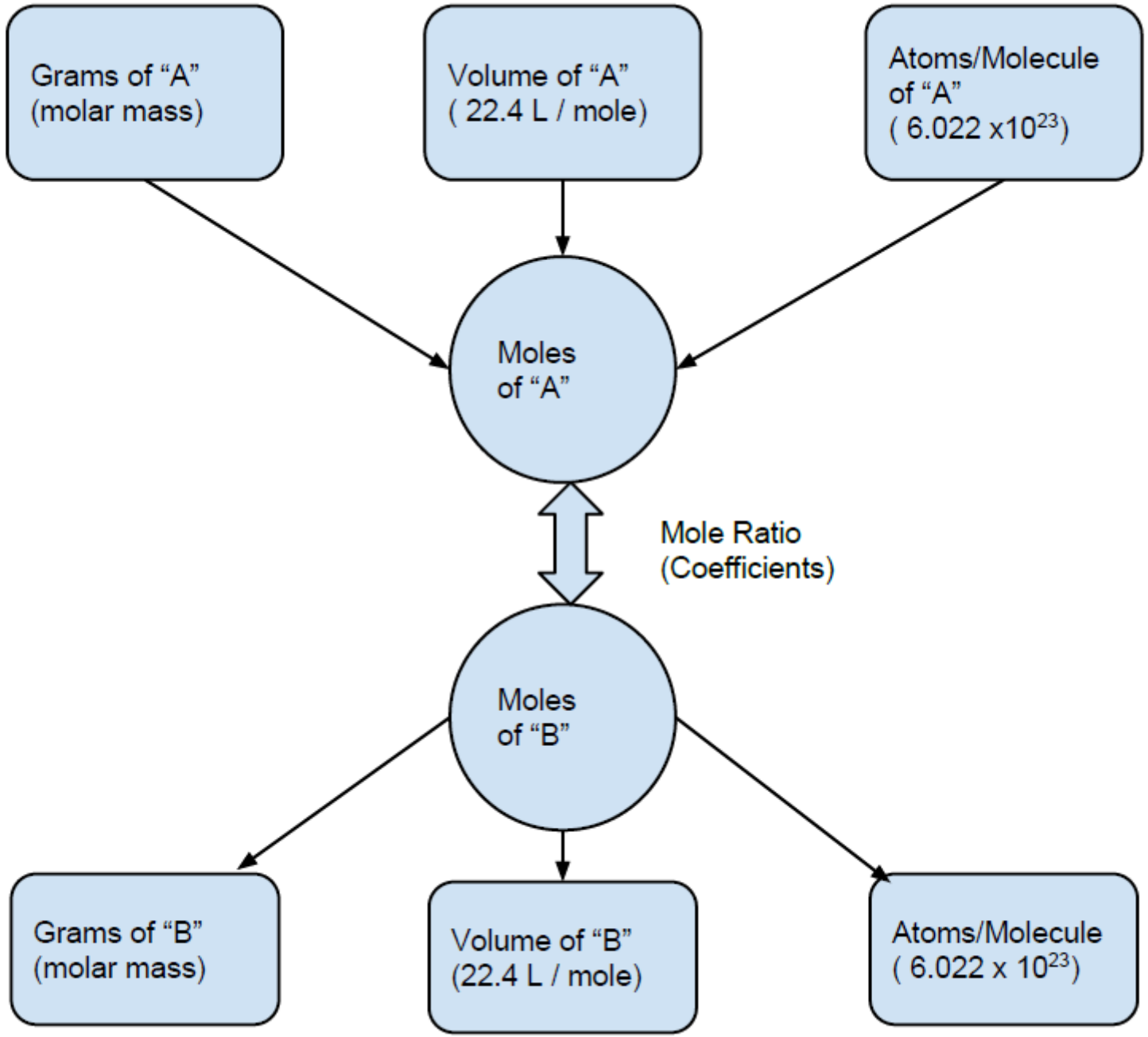
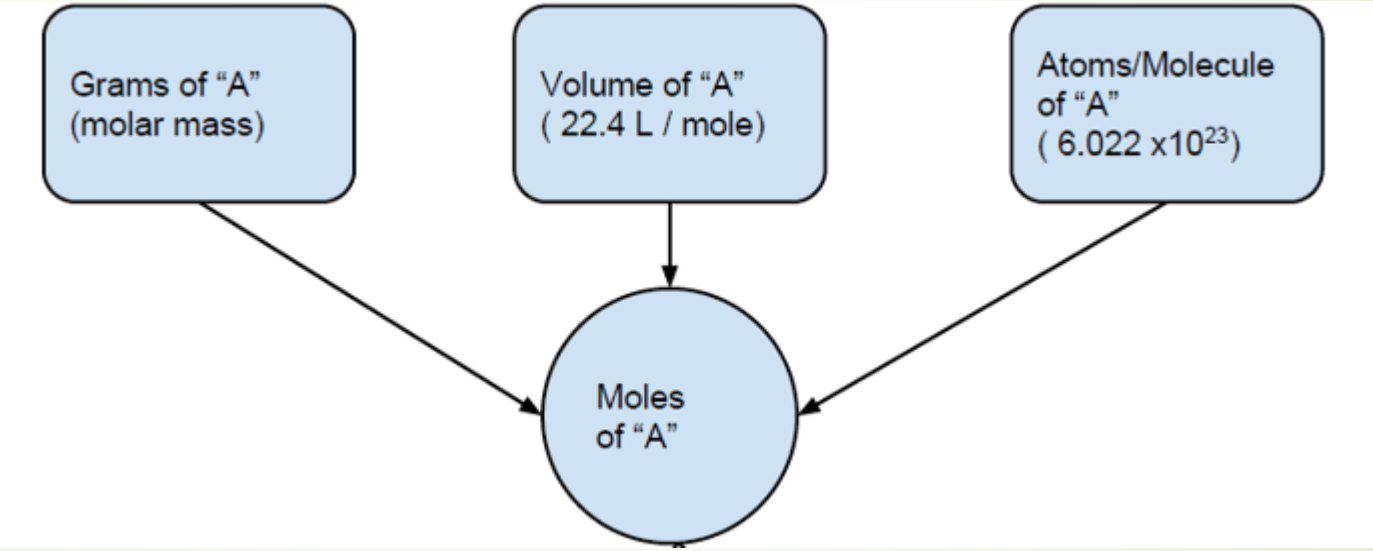
Stoichiometry Help

Determining how much of a specific substance can be created from a specific amount of another substance


Grams of "A"
(molar mass)
Volume of "A"
(22.4 L / mole)
Atoms/Molecule
of "A"
(6.022 x10^23)
Moles
of "A"
Mole Ratio
(Coefficients)
Moles
of "B"
Grams of "B"
(molar mass)
Volume of "B"
(22.4 L / mole)
Atoms/Molecule
(6.022 x 10^23)

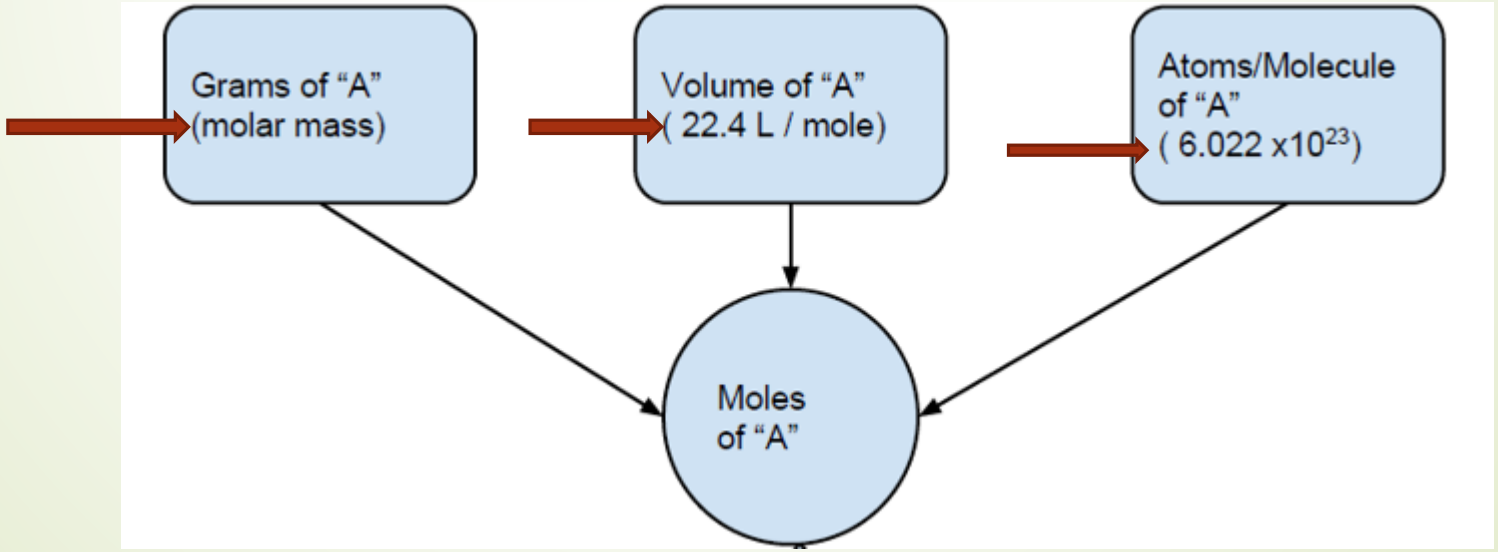
How to use the chart

- You must start with a balanced chemical equation.
- Chose your starting point ("A") by what units you have to start with:



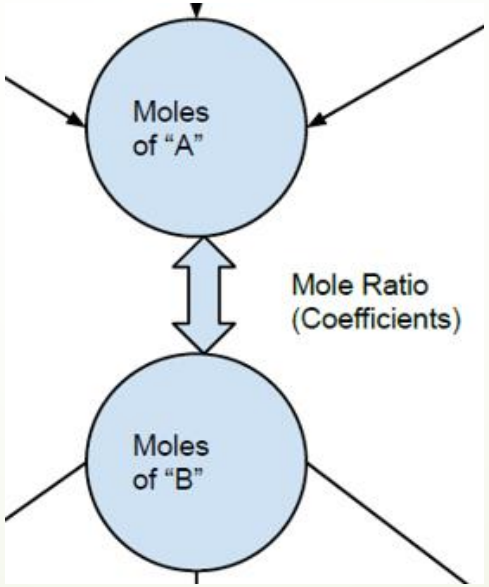
How to use the chart 2

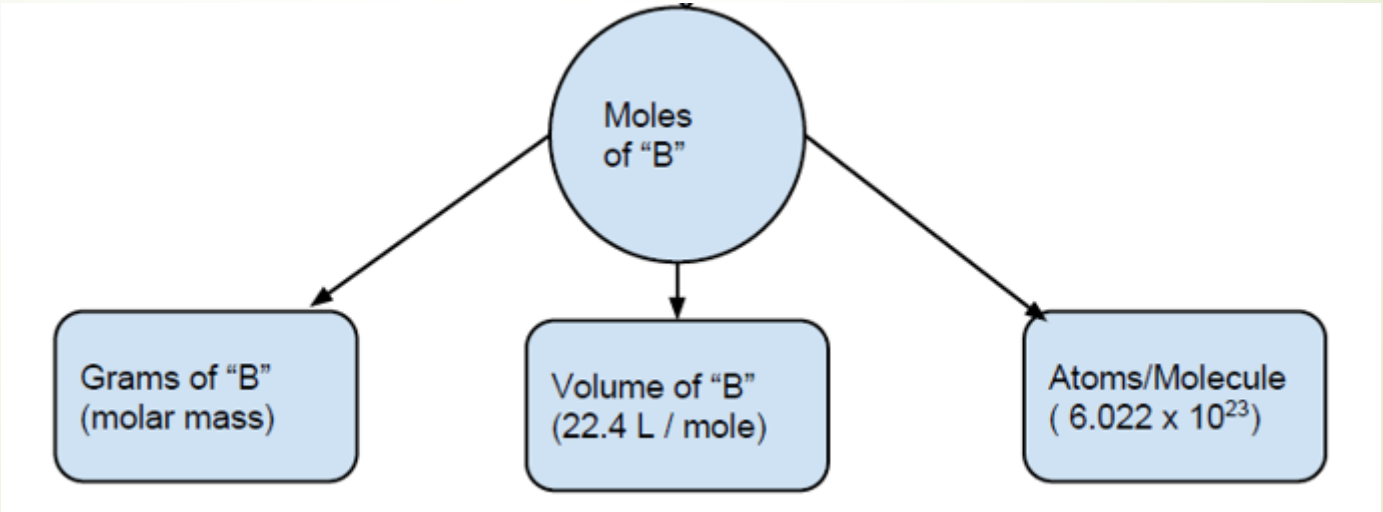
- If you started with any unit other than moles, convert the amount to moles of "A" using dimensional analysis
- You can use the conversion aids listed in parentheses to help.



How to use the chart 3

- Now that you have the moles of "A", set up your next dimensional analysis step to be moles of "B" using the coefficients from the balanced chemical equation as your amounts



How to use the chart 4

- Now that you have the moles of "B" you can convert to any other unit as needed to answer the question.
- Use dimensional analysis and conversion aids are listed in parentheses.